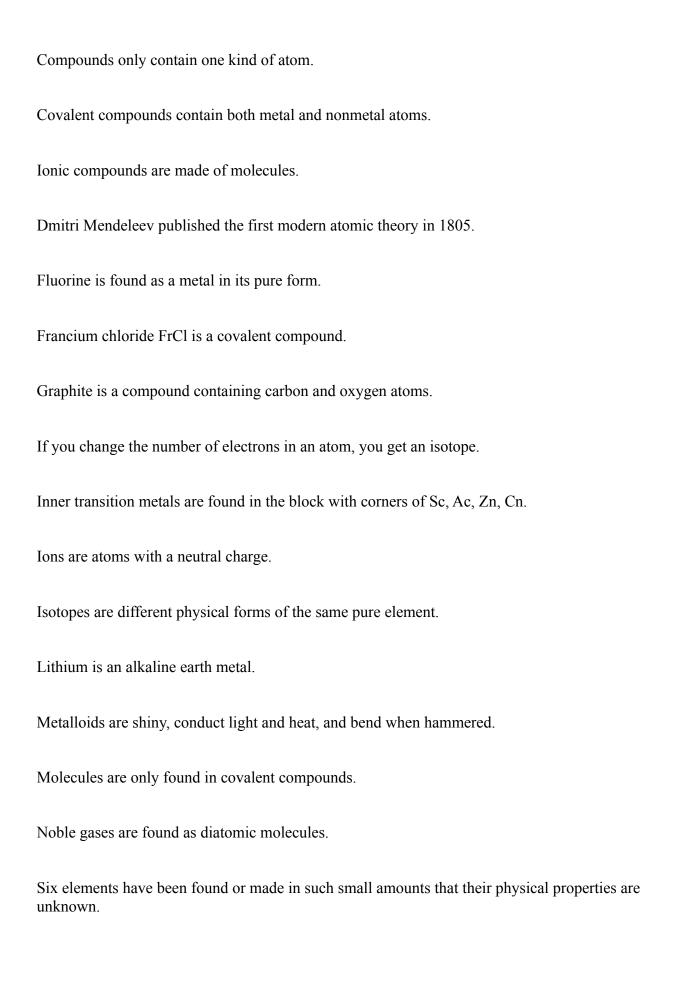
## Review Questions for Unit 1 Chem 1010

Each of the following statements is FALSE. Correct it so that it is true.
A mole of atoms is a certain mass of atoms.
All boron atoms have a mass of 10.81 Daltons.
All of the elements on the left side of the Periodic Table are metals.
All of the elements that are naturally occurring can be found in their pure form in nature.
All of the halogens are gases, and don't react with other elements to form compounds.
Alloys have a set ratio of elements.
Aluminum is in the 4 <sup>th</sup> period in the Periodic Table.
Argon, xenon, krypton, and neon were discovered by Humphry Davy using electrolysis to separate them from their compounds.
Arsenic is in the carbon family.
Astatine is likely to have very similar properties to radon.
Atomic number decreases as you go down a column of the Periodic Table.
Atoms are so small that you can only see them with a microscope.
Chemistry is the study of molecules.
Columns on the Periodic Table are called periods.



The atoms found in bronze have ionic bonds holding them together. The atoms in metals are connected together in diatomic molecules. The most abundant element on the earth is hydrogen. The most common isotope of gallium has a mass of 69 Daltons. The nucleus of an atom contains the protons, neutrons, and electrons. The number of neutrons determines the kind of element an atom is. The symbol for the element mercury is Me. The word "atom" comes from the Greek word for small, and was first suggested by Leucippus. There are 118 naturally occurring elements. There are 9 periods on the Periodic Table of Elements. There are more nonmetals than there are metals. When atoms are connected together, they form an element.